

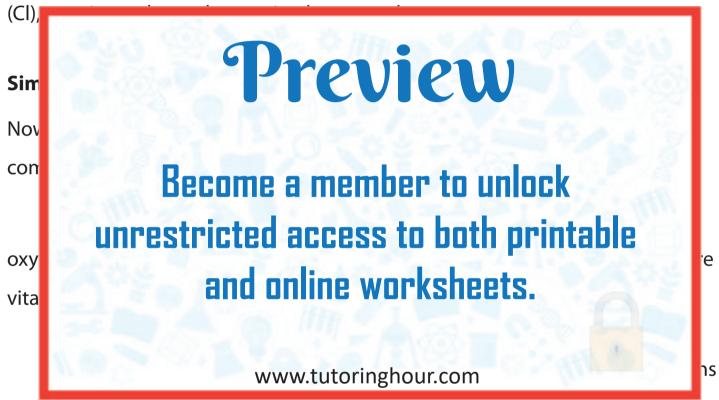
#### **Bonding: Atoms Unite**

Atoms have a strong desire to fill their outermost energy level, and they do so by forming bonds with other atoms. There are two primary ways atoms achieve this:

1. **Covalent Bonds:** In a covalent bond, atoms share electrons to complete their outer shells. For instance, in a molecule of water  $(H_2O)$ , two hydrogen (H)

atoms share their electrons with one oxygen (O) atom. This sharing creates a strong bond that holds the atoms together.

**2. Ionic Bonds:** In an ionic bond, atoms transfer electrons to one another. For example, table salt (NaCl) forms when sodium (Na) donates an electron to chlorine

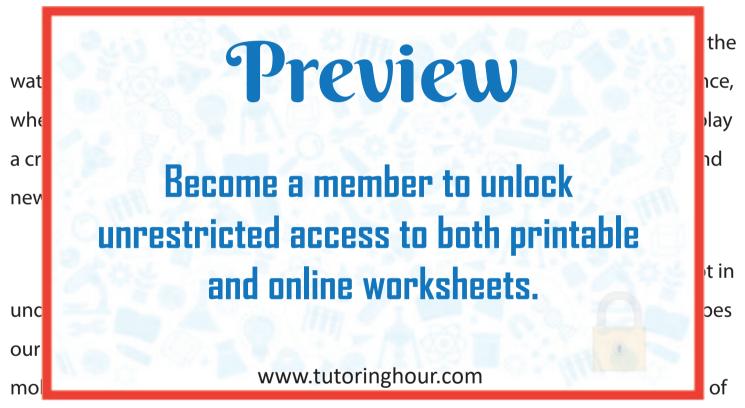


molecules made up of two hydrogen atoms bonded through covalent bonds. Hydrogen is the lightest element in the universe.

• Carbon Dioxide (CO₂): This gas, produced during respiration and combustion, consists of one carbon (C) atom and two oxygen (O) atoms, bonded covalently. It's a critical component in Earth's carbon cycle.

•Ammonia (NH<sub>3</sub>): Ammonia is composed of one nitrogen (N) atom and three hydrogen (H) atoms, bonded covalently. It is commonly used in cleaning products and as a refrigerant.

#### The Importance of Molecules



all the substances we encounter in our daily lives and are the building blocks of the complex chemistry that drives the natural world and our existence within it.

- 1) What is the nucleus of an atom primarily composed of?
  - A) Electrons
  - B) Protons and electrons
  - C) Protons and neutrons
  - D) Neutrons

2)

3)

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- D) It achieves stability by having a complete set of electrons in its outermost shell.
- 4) What is the primary difference between covalent bonds and ionic bonds?

5) What role do molecules play in chemical reactions?

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