

Periodic Table

Identify if the statements are true or false and shade the appropriate circles.

- | | True | False |
|---|-----------------------|-----------------------|
| 1) Atomic size generally decreases as you move from left to right across a period on the periodic table. | <input type="radio"/> | <input type="radio"/> |
| 2) Ionization energy is the energy required to remove an electron from an atom, and it generally decreases as you move from top to bottom within a group on the periodic table. | <input type="radio"/> | <input type="radio"/> |
| 3) Electronegativity values generally increase from left to right across a period on the periodic table. | <input type="radio"/> | <input type="radio"/> |
| 4) The noble gases are the most chemically reactive elements on the periodic table. | <input type="radio"/> | <input type="radio"/> |
| 5) The trend of increasing atomic size from left to right across a period is primarily due to the increasing numbers of protons in the nucleus. | <input type="radio"/> | <input type="radio"/> |
| 6) The trend of decreasing ionization energy from top to bottom within a group is primarily due to the increasing numbers of protons in the nucleus. | <input type="radio"/> | <input type="radio"/> |
| 7) Noble gases, found in Group 18, have the highest ionization energies among all elements. | <input type="radio"/> | <input type="radio"/> |
| 8) Noble gases are the most chemically reactive elements. | <input type="radio"/> | <input type="radio"/> |
| 9) Electronegativity values are generally higher for metals than nonmetals on the periodic table. | <input type="radio"/> | <input type="radio"/> |
| 10) The trend of decreasing atomic size from left to right across a period is primarily due to the increasing numbers of protons in the nucleus. | <input type="radio"/> | <input type="radio"/> |

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